

Poly[μ_2 -chlorido-nonamethyl- μ_3 -nitratotritin(IV)]. Corrigendum

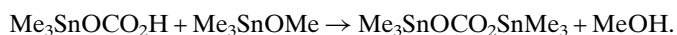
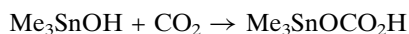
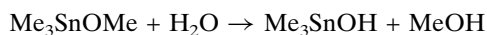
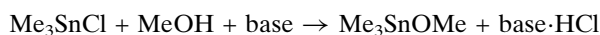
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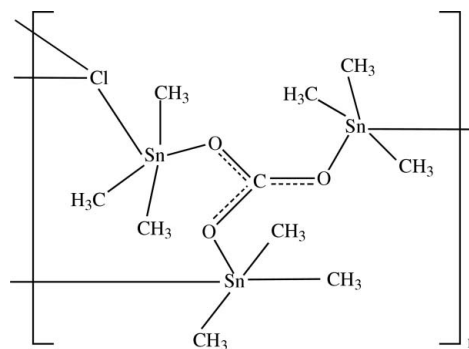
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An error in the original formulation of the title compound in the paper by Sadiq-ur-Rehman, Sherzaman, Ali, Shahzadi & Helliwell [*Acta Cryst.* (2007), **E63**, m2329] is corrected.

The title compound in the paper by Sadiq-ur-Rehman, Sherzaman, Ali, Shahzadi & Helliwell [*Acta Cryst.* (2007), **E63**, m2329] was an unexpected product which seemed to have nitrate coordinated to three Sn atoms. However, it was noticed that the charges do not balance and that it is most likely that the nitrate is in fact a carbonate. Regrettably, there is no material to carry out microanalysis, but a plausible mechanism has been suggested to explain the unexpected formation of the product. Trimethyltin chloride will react with methanol in the presence of a base (4-hydroxypiperidine) to give trimethyltin methoxide, which will rapidly hydrolyze in air to give the hydroxide. Both the methoxide and the hydroxide will react with atmospheric CO₂ to give the carbonate (Bloodworth *et al.*, 1967; Blunden *et al.*, 1984; Sato, 1967).



The carbonate then forms a coordination copolymer with trimethyltin chloride. The name of the title compound is corrected to poly[μ_3 -carbonato- μ_3 -chlorido-nonamethyltritin(IV)], [Sn₃(CH₃)₉(CO₃)Cl] (*M_r* = 586.84).



We thank Professor Alwyn G. Davies (Department of Chemistry, University College London, UK) for providing the mechanism to explain the unexpected formation of the product.

References

- Bloodworth, A. J., Davies, A. G. & Vasishtha, S. V. (1967). *J. Chem. Soc. C*, pp. 1309–1313.
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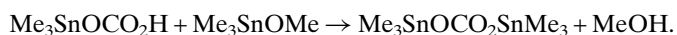
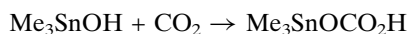
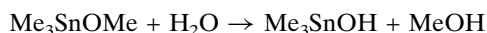
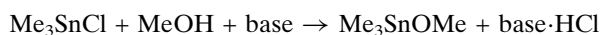
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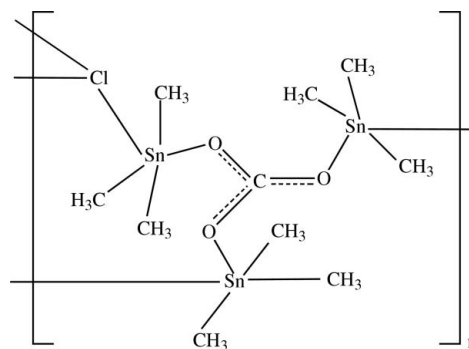
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